CHEMISTRY CHAPTER 6 NOTEPACKET : CHEMICAL BONDS

6.1 Ionic Bonding and 6.3 Naming and Writing Formulas for Ionic Compounds

A. Stable Electron Configurations

1. Electron Dot Diagrams
2. Examples
3. H
4. Be
5. B
6. C
7. N
8. O
9. F
10. Ne

B. Ionic Bonds

1. Transfer of electrons

2. Formation of ions

1. Cation
2. Anion

3. Examples

a. 23Na+

b. 19F-

c. 24Mg2+

d. 15N3-

4. Formation of Ionic Bonds

5. Ionization Energy

C. Ionic Compounds

1. Crystal Lattices
2. Properties of Ionic Compounds

D. Describing Ionic Compounds

-Type I

-Type II

Polyatomic Ions

1. Type I Binary Ionic Compounds-Naming
2. Type II Binary Ionic Compounds-Naming
3. Ionic Compounds Containing Polyatomic Ions-Naming

1. Formula Writing for Ionic Compounds

6.2 Covalent Bonding and 6.3 Naming and Writing Formulas for Type III Compounds

A. Covalent Bonds

1. Sharing Electrons
2. Molecules Of Elements
3. Multiple Covalent Bonds

B. Unequal Sharing of Electrons

1. Polar Covalent Bonds

2.Polar and Nonpolar Molecules

C. Attraction Between Molecules

D. Describing Molecular Compounds

1. Type III Compounds -Naming
2. Type III compounds- Formula Writing
3. Properties of Molecular Compounds

6.4 The Structure of Metals

1. Metallic Bonds
2. Explaining Properties of Metals
3. Alloys
4. Copper Alloys
5. Steel Alloys
6. Other Alloys